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Transport and transformation processes are key for determining how humans and other organisms are exposed to chemicals. These processes are largely controlled by the chemicals’ physical-chemical

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Written by one of the most prolific and well-respected chemical engineers in the industry, this is the most comprehensive and thorough volume ever written on the thermodynamic properties of hydrocarbons and chemicals. This volume covers the spectrum, including chapters on the heat capacity and entropy of gas, solids and liquids, the entropy of formation, and many other topics. The design of heat exchangers and other equipment for heating or cooling substances to temperatures necessary in process applications requires knowledge of heat capacity, covered in the first portion of the book. The heat effects of chemical reactions are ascertained from enthalpy of formation. Other chapters cover the Helmholtz energy of formation and internal energy of formation, which is useful in modeling and ascertaining the energy of explosions. This coverage greatly exceeds the coverage of any other book and makes The Yaws Handbook of Thermodynamic Properties of Hydrocarbons and Chemicals a must-have for anyone working in the fields of chemical engineering, process engineering, refining and chemistry.

This book is an excellent companion to Chemical Thermodynamics: Principles and Applications. Together they make a complete reference set for the practicing scientist. This volume extends the range of topics and applications to ones that are not usually covered in a beginning thermodynamics text. In a sense, the book covers a "middle ground" between the basic principles developed in a beginning thermodynamics textbook, and the very specialized applications that are a part of an ongoing research project. As such, it could prove invaluable to the practicing scientist who needs to apply thermodynamic relationships to aid in the understanding of the chemical process under consideration. The writing style in this volume remains informal, but more technical than in Principles and Applications. It starts with Chapter 11, which summarizes the thermodynamic relationships developed in this earlier volume. For those who want or need more detail, references are given to the sections in Principles and Applications where one could go to learn more about the development, limitations, and conditions where these equations apply. This is the only place where Advanced Applications ties back to the previous volume. Chapter 11 can serve as a review of the fundamental thermodynamic equations that are necessary for the more sophisticated applications described in the remainder of this book. This may be all that is necessary for the practicing scientist who has been away from the field for some time and needs some review. The remainder of this book applies thermodynamics to the description of a variety of problems. The topics covered are those that are probably of the most fundamental and broadest interest. Throughout the book, examples of "real" systems are used as much as possible. This is in contrast to many books where "generic" examples are used almost exclusively. A complete set of references to all sources of data and to supplementary reading sources is included. Problems are given at the end of each chapter. This makes the book ideally suited for use as a textbook in an advanced topics course in chemical thermodynamics. An excellent review of thermodynamic principles and mathematical relationships along with references to the relevant sections in Principles and Applications where these equations are developed Applications of thermodynamics in a wide variety of chemical processes, including phase equilibria, chemical equilibrium, properties of mixtures, and surface chemistry Case-study approach to demonstrate the application of thermodynamics to biochemical, geochemical, and industrial processes Applications at the "cutting edge" of thermodynamics Examples and problems to assist in learning Includes a complete set of references to all literature sources

The American edition of this handbook contains concise informa tion on the basic physical properties of the elements and on their chem ical characteristics. In general, the data selected for inclusion in the handbook are those which either agree well with calculated data (in those cases where calculations could be carried out) or satisfy various correlations, particularly those based on concepts of the distribution of valence electrons of isolated atoms in the formation of a condensed state, as electrons localized at atomic ions in the form of energetically stable configurations, and as nonlocalized electrons. The Russian edition was published in the USSR in 1965, and new or previously omitted data have been added to all the sections of the present edition. In addition, the authors have considered it necessary to include a series of new sections. Thus, a new table has been included, "Electronic Configurations and Ground States of Free Atoms and Their Ions," since, in the ionization of some atoms (particularly for transition metals), the electrons are not always abstracted from the outer shell, and, consequently, calculation of the ground state (electron energy level) using the usual vector model does not give a direct result. The ground states are obtained experimentally and the table contains the corresponding data on the configurations and states of triply-ionized atoms (which is usually sufficient).

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This volume is a compilation of data on the properties of glasses. The authors have critically examined and correlated the most reliable data on the properties of multicomponent commercial silicate glasses, vitreous silica, and binary and ternary laboratory glasses. Thermodynamic, thermal, mechanical, electrical, and transport properties are covered. Measurement methods and appropriate theories are also discussed.

This practical handbook features an overview of the importance of physical properties and thermodynamics; and the use of thermo-dynamics to predict the extent of reaction in proposed new chem-ical combinations. The use of special types of data and pre-diction methods to develop flowsheets for probing projects; and sources of critically evaluated data, dividing the published works into three categories depending on quality are given. Methods of doing one's own critical evaluation of literature, a list of known North American contract experimentalists with the types of data mea-sured by each, methods for measuring equilibrium data, and ther-modynamic concepts to carry out process opti-mization are also featured.

Expertise in electrolyte systems has become increasingly important in traditional CPI operations, as well as in oil/gas exploration and production. This book is the source for predicting electrolyte systems behavior, an indispensable "do-it-yourself" guide, with a blueprint for formulating predictive mathematical electrolyte models, recommended tabular values to use in these models, and annotated bibliographies. The final chapter is a general recipe for formulating complete predictive models for electrolytes, along with a series of worked illustrative examples. It can serve as a useful research and application tool for the practicing process engineer, and as a textbook for the chemical engineering student.