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Problems | How to Pass

Chemistry

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**Chemical reactions
and stoichiometry |**

Chemistry | Khan

Academy ~~Limiting~~

~~Reactant Practice~~

~~Problems~~ Stoichiometry

Basic Introduction,

Mole to Mole, Grams to

Grams, Mole Ratio

Practice Problems

Stoichiometry - Limiting

& Excess Reactant,

Theoretical &

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Percent Yield -

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Chemistry #6 How To

Solve Stoichiometry

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Chemistry How to

Calculate Percent Yield

and Theoretical Yield

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Stoichiometry Made

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Tutorial Part 1 Mole

~~Conversions Made~~

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~~Number Method~~

Limiting Reactant

Practice Problem

(Advanced)

Limiting Reactant

Practice Problem

Limiting Reagents and

Percent Yield

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~~Converting Grams to~~
~~Grams~~ *Easiest way to*
solve limiting reagent
problems - ABCs of
limiting reagent

~~Hydrocarbon Power!:~~

~~Crash Course Chemistry~~

~~#40~~ *Introduction to*

Oxidation Reduction

(Redox) Reactions

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Part 1 Thermochemistry

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Stoichiometry
Equations \u0026amp;

Formulas - Lecture

Review \u0026amp; Practice

Problems HOW TO

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Table, Elements, Metric

System \u0026amp; Unit

Conversion Concept of

Mole - Part 1 | Atoms

and Molecules | Don't

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11.1 Defining

Stoichiometry Vocab •

stoichiometry • mole

ratio Concepts Using

Balanced Equations •

Number of Atoms •

Number of Molecules •

Number of Moles

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Stoichiometry Section

111 What Is

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STOICHIOMETRY

Mole-a unit of

measurement that

expresses the amount of

atoms, molecules or

some other unit. The

number of items in one

mole is commonly

referred to Avogadro's

number which equals

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- ratio Concepts Using
- Balanced Equations •
- Number of Atoms •
- Number of Molecules •
- Number of Moles •
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Monday (1/27): In-
class: Naming, Formula,
and Balancing Review;
HW: Finish Review,
Watch Stoichiometry
Lecture Video and

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Complete Guided Notes

(p 1-3) Due Tomorrow

Tuesday (1/28): In-

class: Note Check,

Funsheets A-E; HW:

Study for Quiz Friday

Wednesday (1/29): In-

class: Funsheets A-E;

HW: Watch Percent

Yield and Limiting

Reactant Lecture Video

and Complete Guided

Notes (p 4-6) Due

Monday ...

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Guide - Chapter 11 –

Stoichiometry Section

11.1 What is

stoichiometry? 1. true 2.

true 3. false 4. true 5.

true 6. 2, 2, 64.10 7. 3,

3, 96.00 8. 2, 2, 88.02 9.

4, 4, 72.08 10. methanol

and oxygen gas 11.

carbon dioxide and

water 12. 160.10 g 13.

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160.10 g 14. They are equal. 15. A mole ratio is a ratio between ...

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balanced equation. For each statement below, write true or false. _____
_____ true _____ 1.

The study of the quantitative relationships between the amounts of reactants used and the amounts of products formed by a chemical reaction is called stoichiometry. _____
_____ true _____

2.

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